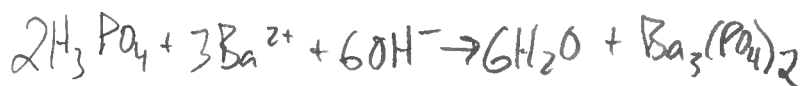
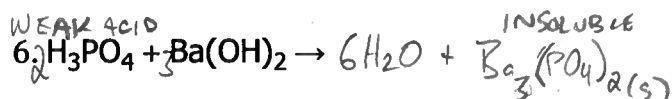
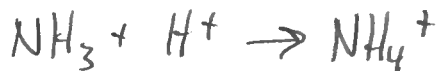
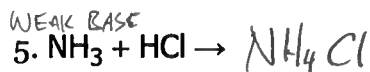
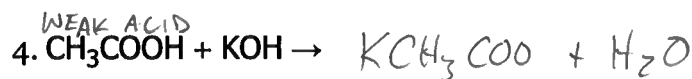
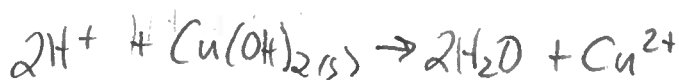
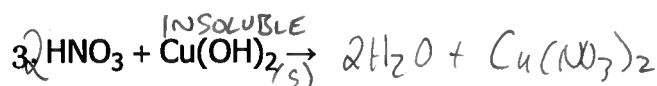
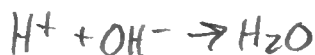
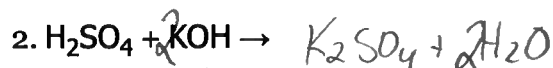
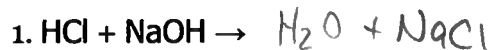


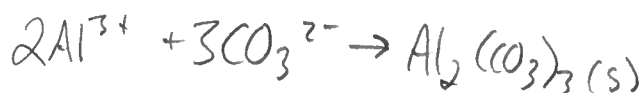
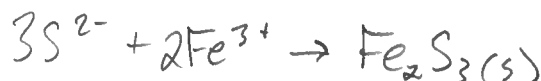
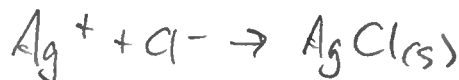
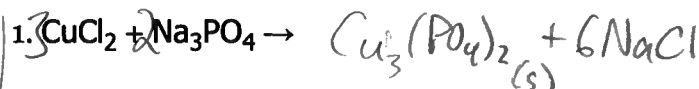
Net Ionic Equations

Predict products for the following chemical reactions. Write **molecular equations** and balance them. Write **total ionic equations** and **net ionic equations**. Insoluble substances, non-electrolytes, and weak electrolytes remain in molecular form when writing ionic equations. Note especially that care must be taken in determining the ionic composition of compounds. For example, sodium phosphate (Na_3PO_4) forms ions in solution to produce three Na^+ ions and one PO_4^{3-} ion. It does not produce an ion with this formula: Na_3^+ .

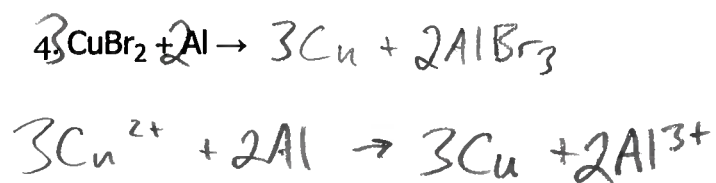
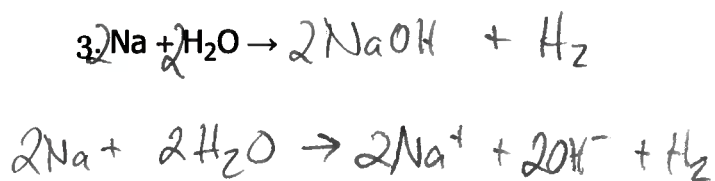
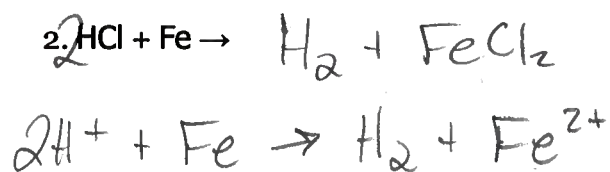
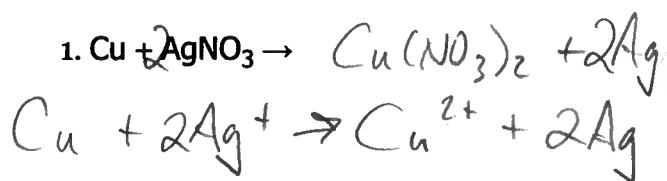
Acid-Base



Precipitation



Single Replacement



Gas Formation

